

The Crystal Structure of Ryanodol *p*-Bromobenzyl Ether*

BY SURENDRA NATH SRIVASTAVA† AND MARIA PRZYBYLSKA

Biochemistry Laboratory, National Research Council of Canada, Ottawa, Canada

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The crystals of ryanodol-*p*-bromobenzyl ether, $C_{27}H_{37}O_8Br$, belong to the space group $P2_12_12_1$ and the unit-cell dimensions are: $a=17.65 \pm 0.03$, $b=14.43 \pm 0.02$ and $c=10.11 \pm 0.02$ Å. The structure was solved by the heavy-atom method from three-dimensional X-ray intensity data collected photographically. The atomic parameters were refined by block-diagonal least squares to the R value of 0.090. The results confirm the structure and relative stereochemistry independently deduced by chemical means by K. Wiesner, except for the configuration at one of the eleven asymmetric centres. Only one hydrogen bond (intermolecular) was found. All other hydrogen atoms of OH groups make intra- and intermolecular contacts with oxygen atoms which do not deviate from the sum of the van der Waals radii.

Introduction

The insecticidal principle ryanodine has been isolated from the stems and roots of *Ryania speciosa* Vahl. by Folkers and his co-workers (Rogers, Koniuszy, Shavel & Folkers, 1948).

A chemical investigation of ryanodine was undertaken by K. Wiesner and his collaborators in 1951, and they established that this compound is an ester of pyrrole- α -carboxylic acid and of the fully saturated polyalcohol ryanodol, $C_{20}H_{32}O_8$. Fig. 1 shows the schematic formula as established by this X-ray investigation and the relationship between ryanodine, ryanodol and ryanodol *p*-bromobenzyl ether. The extensive degradation studies led first to the elucidation of the structure and the determination of the absolute configuration of anhydroryanodine, the dehydration product of ryanodine (Babin, Forrest, Valenta & Wiesner, 1962; Wiesner, 1963) and later to the proposals for the structures of ryanodol and ryanodine (Wiesner, Valenta & Findlay, 1967; Wiesner, 1968).

This paper reports the crystal structure determination of ryanodol *p*-bromobenzyl ether, $C_{27}H_{37}O_8Br$, and confirms the molecular structures for this and related compounds proposed by Wiesner *et al.*, except for their assignment of configuration at the carbon atom carrying the isopropyl group.

The absolute configuration of ryanodol *p*-bromobenzyl ether has not been determined by X-rays, but the skeleton in Fig. 1 and throughout this report conforms with the absolute configuration derived by Wiesner (1963) for anhydroryanodine.

A preliminary account of this work was published earlier (Srivastava & Przybylska, 1968).

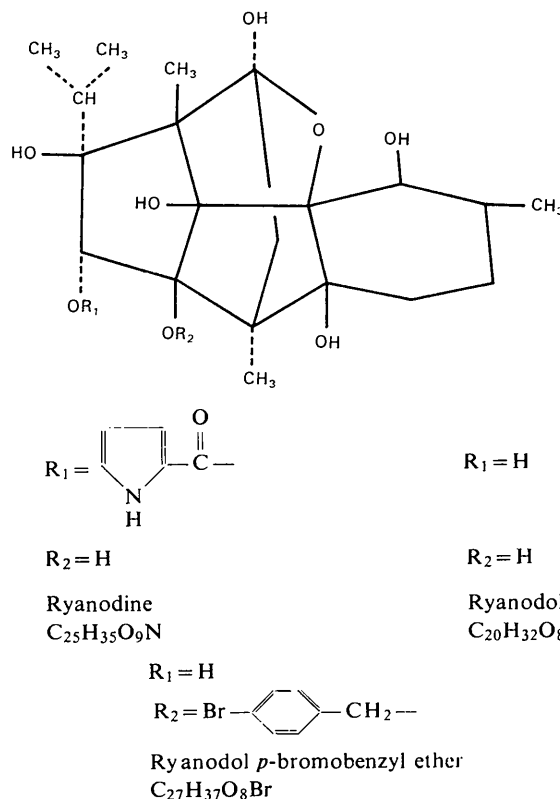


Fig. 1. Schematic formula of ryanodol *p*-bromobenzyl ether and its relationship to ryanodine and ryanodol.

Crystal data

$C_{27}H_{37}O_8Br$, M.W. 569.5

Orthorhombic; the systematic absences determined the space group as $P2_12_12_1$ (D_2^7). There are four molecules per unit cell of dimensions: $a=17.65 \pm 0.03$, $b=14.43 \pm 0.02$, $c=10.11 \pm 0.02$ Å.

$D_m = 1.463$ g.cm $^{-3}$ (by flotation in a mixture of carbon

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† National Research Council of Canada Postdoctorate Fellow, Division of Pure Chemistry, 1965–1967. Present address: Physics Department, University of Allahabad, Allahabad, India.

tetrachloride and toluene) and $D_c=1.462 \text{ g.cm}^{-3}$. $F(000)=1192$. $\mu=28.1 \text{ cm}^{-1}$. Results of the chemical analysis obtained from Dr Wiesner:

Calculated: C, 56.94; H, 6.55; O, 22.48; Br, 14.05%.
Found: C, 56.79; H, 6.34; O, 22.73; Br, 14.01%.

The crystals from Dr Wiesner were obtained by evaporating slowly a solution of the compound in a mixture of methanol and chloroform. Various attempts in our laboratory to increase the crystals to the size suitable for X-ray work proved unsuccessful, until Dr W. Mechlinski used a minimum amount of ethyl acetate to dissolve the crystals and then added petroleum spirit until the solution became slightly opaque.

The lattice constants were determined from precession photographs (Cu $K\alpha=1.5418 \text{ \AA}$). Sets of multiple-film, equi-inclination Weissenberg photographs were taken for the crystal rotating about the *b* axis (layer $h0l$ through $h4l$) and about the *c* axis (layer $hk0$ through $hk8$). A total of 2406 independent reflexions were measured by visual comparison with the standard intensity strip prepared with a crystal of the same compound. The reflexions, which were measured twice, were averaged and the absolute scale and overall temperature factors were evaluated by Wilson's (1942) method. The crystals were cut so that they were cylindrical in shape. Their diameters ranged between 0.2 and 0.3 mm and absorption corrections were not applied.

Solution of the structure and its refinement

The position of the bromine atom was derived from the Harker vectors of a three-dimensional Patterson synthesis which was sharpened by the $1/f_{\text{Br}}^2$ function. The first electron density maps calculated with the phases based only on the contribution of the heavy atom revealed the location of ten light atoms, which were treated as carbon atoms. Two subsequent Fourier cycles gave eight and nine additional atoms bringing the total of light atoms to 27 and the *R* value to 0.32. The remaining eight atoms of the molecule were obtained from three-dimensional difference maps. The refinement of the parameters by least-squares was then undertaken and two additional $\rho_o - \rho_c$ syntheses were carried out in order to distinguish the oxygen atoms from carbon atoms. A value for $R=0.17$ was obtained when the molecular structure was solved. Several cycles of block-diagonal least-squares, with the introduction of anisotropic refinement for all atoms, led to a considerable improvement between the calculated and observed structure factors, giving a reliability index of 0.115.

At this stage the hydrogen atoms were located. First the positions of sixteen hydrogen atoms, those attached to the benzene ring and to di- and tri-substituted carbon atoms, were computed. They were included in the calculation of structure factors used for the evaluation of the last difference synthesis. The re-

maining twenty-one hydrogen atoms were located on these maps without any difficulty and only the position of a hydrogen atom attached to O(30) had to be adjusted. It was found to be too close to the oxygen atom, and it was moved by about 0.3 Å along the O-H line. In some methyl groups one hydrogen atom did not show up as clearly as the other two, but it was possible to derive their coordinates using the positions of the acceptable hydrogen atoms.

Four additional least-squares cycles were carried out until the shifts for two-thirds of the parameters were less than $\frac{1}{3}$ of their corresponding standard deviations and for the remaining coordinates the shifts were larger, but did not exceed $\frac{1}{3}$ of their e.s.d.'s.

The final value of $R=0.090$ was obtained excluding the unobserved reflexions. The atomic coordinates and their estimated standard deviations of the non-hydrogen atoms are given in Table 1 and those of the hydrogen atoms are presented in Table 2. The positions of the hydrogen atoms were not refined and the temperature factor $B=5.0 \text{ \AA}^2$ was used in the calculation of their contribution. The final vibrational parameters and the estimated standard deviations of the non-hydrogen atoms are listed in Table 3, together with the values of the principal radii of the vibration ellipsoids.

Table 1. Final fractional coordinates and their estimated standard deviations

| | <i>x/a</i> | <i>y/b</i> | <i>z/c</i> |
|-------|-------------|------------|--------------|
| Br | -0.1042 (1) | 0.6888 (1) | -0.3019 (1) |
| C(1) | 0.1794 (5) | 0.2413 (6) | 0.2842 (10) |
| C(2) | 0.0943 (5) | 0.2586 (5) | 0.2418 (9) |
| C(3) | 0.0803 (4) | 0.3643 (5) | 0.2567 (9) |
| C(4) | 0.0457 (5) | 0.3956 (5) | 0.3896 (9) |
| C(5) | 0.0663 (4) | 0.5003 (5) | 0.3924 (9) |
| C(6) | 0.0533 (6) | 0.5574 (6) | 0.5126 (9) |
| C(7) | 0.0839 (6) | 0.6554 (6) | 0.4868 (11) |
| C(8) | 0.1649 (5) | 0.6599 (6) | 0.4359 (11) |
| C(9) | 0.1836 (4) | 0.5891 (5) | 0.3253 (8) |
| C(10) | 0.1517 (4) | 0.4955 (5) | 0.3604 (8) |
| O(11) | 0.1901 (3) | 0.4630 (4) | 0.4770 (6) |
| C(12) | 0.1788 (5) | 0.3630 (5) | 0.4787 (9) |
| C(13) | 0.2061 (5) | 0.3376 (5) | 0.3363 (10) |
| C(14) | 0.1580 (4) | 0.4114 (6) | 0.2617 (8) |
| C(15) | 0.0940 (5) | 0.3479 (6) | 0.5005 (9) |
| C(16) | 0.1899 (6) | 0.1531 (6) | 0.3721 (12) |
| C(17) | 0.1680 (7) | 0.0663 (6) | 0.2952 (15) |
| C(18) | 0.2698 (8) | 0.1418 (9) | 0.4233 (15) |
| O(19) | 0.2201 (4) | 0.2266 (5) | 0.1638 (7) |
| O(20) | 0.0387 (4) | 0.2066 (4) | 0.3161 (7) |
| O(21) | 0.0292 (3) | 0.3963 (4) | 0.1569 (6) |
| C(22) | 0.0403 (6) | 0.3750 (6) | 0.0206 (8) |
| C(23) | 0.0052 (5) | 0.4500 (6) | -0.0580 (9) |
| C(24) | -0.0574 (7) | 0.4334 (7) | -0.1352 (11) |
| C(25) | -0.0916 (7) | 0.5049 (9) | -0.2069 (13) |
| C(26) | -0.0602 (6) | 0.5914 (8) | -0.2006 (11) |
| C(27) | 0.0024 (6) | 0.6085 (7) | -0.1254 (11) |
| C(28) | 0.0345 (6) | 0.5374 (6) | -0.0571 (11) |
| C(29) | -0.0377 (5) | 0.3801 (7) | 0.4120 (10) |
| O(30) | 0.0281 (3) | 0.5483 (4) | 0.2877 (6) |
| C(31) | 0.1825 (7) | 0.7584 (6) | 0.3865 (13) |
| O(32) | 0.2640 (3) | 0.5830 (4) | 0.3062 (7) |
| O(33) | 0.2236 (3) | 0.3258 (4) | 0.5747 (6) |
| C(34) | 0.2922 (5) | 0.3535 (7) | 0.3284 (10) |
| O(35) | 0.1857 (3) | 0.4369 (4) | 0.1343 (6) |

Table 2. Fractional coordinates of the hydrogen atoms

The number of the parent atom is obtained by omitting the last digit. The atoms marked with an asterisk were found experimentally.

| | <i>x/a</i> | <i>y/b</i> | <i>z/c</i> |
|---------|------------|------------|------------|
| H(21) | 0.086 | 0.237 | 0.142 |
| H(61) | 0.082 | 0.526 | 0.595 |
| H(62) | -0.006 | 0.561 | 0.534 |
| H(71) | 0.084 | 0.694 | 0.578 |
| H(72) | 0.050 | 0.689 | 0.415 |
| H(81) | 0.202 | 0.647 | 0.518 |
| H(91) | 0.158 | 0.613 | 0.236 |
| H(151) | 0.080 | 0.375 | 0.595 |
| H(152) | 0.084 | 0.274 | 0.499 |
| H(161) | 0.157 | 0.160 | 0.460 |
| H(171)* | 0.172 | 0.007 | 0.356 |
| H(172)* | 0.207 | 0.057 | 0.210 |
| H(173)* | 0.110 | 0.072 | 0.258 |
| H(181)* | 0.285 | 0.199 | 0.481 |
| H(182)* | 0.309 | 0.135 | 0.342 |
| H(183)* | 0.272 | 0.080 | 0.487 |
| H(191)* | 0.217 | 0.270 | 0.085 |
| H(201)* | 0.085 | 0.155 | 0.303 |
| H(221) | 0.013 | 0.311 | -0.002 |
| H(222) | 0.099 | 0.370 | 0.000 |
| H(241) | -0.081 | 0.363 | -0.144 |
| H(251) | -0.144 | 0.490 | -0.262 |
| H(271) | 0.026 | 0.677 | -0.123 |

Table 2 (cont.)

| | <i>x/a</i> | <i>y/b</i> | <i>z/c</i> |
|---------|------------|------------|------------|
| H(281) | 0.087 | 0.551 | -0.002 |
| H(291)* | -0.059 | 0.319 | 0.363 |
| H(292)* | -0.070 | 0.439 | 0.374 |
| H(293)* | -0.048 | 0.371 | 0.518 |
| H(301)* | -0.028 | 0.547 | 0.287 |
| H(311)* | 0.180 | 0.809 | 0.468 |
| H(312)* | 0.238 | 0.760 | 0.338 |
| H(313)* | 0.143 | 0.779 | 0.309 |
| H(321)* | 0.293 | 0.557 | 0.382 |
| H(331)* | 0.215 | 0.360 | 0.659 |
| H(341)* | 0.308 | 0.423 | 0.297 |
| H(342)* | 0.317 | 0.303 | 0.259 |
| H(343)* | 0.317 | 0.341 | 0.427 |
| H(351)* | 0.240 | 0.444 | 0.142 |

The scattering factors for the non-hydrogen atoms were those of Hanson, Herman, Lea & Skillman (1964) and for the hydrogen atoms those of Stewart, Davidson & Simpson (1965).

The observed and calculated structure factors are listed in Table 4. The structure factors for the unobserved reflexions were calculated, but they were excluded from the refinement procedure and from general sums. The agreement summary is given in Table 5. Two

Table 3. Vibration tensor components and their *e.s.d.*'s ($\text{\AA}^2 \times 10^3$)

Temperature factor = $\exp \{-2\pi^2(U_{11}a^2h^2 + U_{22}b^2k^2 + U_{33}c^2l^2 + 2U_{23}b^*c^*kl + 2U_{13}a^*c^*hl + 2U_{12}a^*b^*hk)\}$.

B_i (\AA^2) are the principal radii of the vibration ellipsoids.

| | U_{11} | U_{22} | U_{33} | $2U_{23}$ | $2U_{13}$ | $2U_{12}$ | B_1 | B_2 | B_3 |
|-------|----------|----------|----------|-----------|-----------|-----------|-------|-------|-------|
| Br | 113 (1) | 95 (1) | 64 (1) | 32 (1) | 8 (2) | 77 (2) | 11.51 | 5.76 | 4.20 |
| C(1) | 50 (4) | 47 (4) | 59 (6) | -22 (8) | 11 (8) | 4 (7) | 5.23 | 4.01 | 3.02 |
| C(2) | 51 (4) | 35 (3) | 38 (4) | -19 (6) | -6 (7) | -15 (6) | 4.24 | 3.56 | 1.94 |
| C(3) | 45 (4) | 41 (4) | 35 (4) | -6 (7) | 5 (6) | -6 (6) | 3.78 | 3.15 | 2.69 |
| C(4) | 45 (4) | 40 (3) | 44 (5) | 3 (7) | 15 (7) | 3 (6) | 4.10 | 3.16 | 2.88 |
| C(5) | 42 (4) | 39 (3) | 42 (5) | 14 (7) | 0 (7) | 14 (6) | 3.97 | 3.30 | 2.43 |
| C(6) | 67 (5) | 45 (4) | 43 (5) | -15 (8) | 14 (9) | 5 (8) | 5.42 | 4.07 | 2.77 |
| C(7) | 69 (6) | 43 (4) | 55 (6) | -13 (8) | -5 (9) | 16 (8) | 5.72 | 4.40 | 3.09 |
| C(8) | 52 (4) | 42 (4) | 63 (6) | -12 (8) | -10 (9) | 1 (7) | 5.26 | 3.95 | 3.22 |
| C(9) | 48 (4) | 35 (3) | 38 (5) | -12 (6) | 2 (7) | 0 (6) | 3.82 | 3.36 | 2.40 |
| C(10) | 46 (4) | 38 (3) | 33 (4) | 3 (7) | -2 (7) | 3 (6) | 3.69 | 3.02 | 2.59 |
| O(11) | 48 (3) | 37 (2) | 43 (3) | -2 (5) | -16 (5) | 2 (4) | 4.27 | 2.94 | 2.90 |
| C(12) | 49 (4) | 37 (3) | 39 (5) | -10 (7) | -8 (7) | -3 (6) | 4.00 | 3.37 | 2.57 |
| C(13) | 45 (4) | 38 (3) | 57 (6) | 6 (7) | 3 (8) | -4 (6) | 4.54 | 3.60 | 2.95 |
| C(14) | 42 (3) | 47 (4) | 33 (4) | -16 (7) | 4 (6) | -19 (6) | 4.47 | 2.87 | 2.28 |
| C(15) | 56 (4) | 44 (4) | 37 (4) | 2 (7) | -1 (7) | -1 (7) | 4.43 | 3.51 | 2.89 |
| C(16) | 60 (5) | 45 (4) | 81 (7) | 9 (9) | 9 (10) | 6 (8) | 6.50 | 4.65 | 3.47 |
| C(17) | 80 (7) | 37 (4) | 115 (10) | 2 (11) | -29 (15) | 6 (9) | 9.51 | 5.97 | 2.94 |
| C(18) | 92 (8) | 68 (6) | 93 (9) | -23 (14) | -45 (15) | 39 (13) | 9.80 | 5.74 | 4.53 |
| O(19) | 68 (4) | 67 (4) | 57 (4) | -37 (7) | 27 (7) | 2 (7) | 6.73 | 5.41 | 3.05 |
| O(20) | 59 (3) | 49 (3) | 58 (4) | -11 (6) | 11 (6) | -24 (5) | 5.63 | 4.23 | 3.23 |
| O(21) | 51 (3) | 50 (3) | 39 (3) | 4 (5) | -5 (5) | -2 (5) | 4.10 | 3.90 | 3.01 |
| C(22) | 73 (5) | 42 (4) | 29 (4) | -10 (7) | -11 (8) | -20 (8) | 6.00 | 3.31 | 2.06 |
| C(23) | 56 (5) | 50 (4) | 37 (5) | -2 (8) | -14 (8) | -16 (7) | 4.96 | 3.70 | 2.70 |
| C(24) | 88 (7) | 62 (5) | 54 (6) | 7 (10) | -40 (11) | -43 (11) | 8.43 | 4.33 | 3.33 |
| C(25) | 76 (6) | 82 (7) | 67 (7) | 22 (13) | -45 (12) | -18 (12) | 8.17 | 5.74 | 3.83 |
| C(26) | 67 (6) | 85 (7) | 39 (5) | 20 (11) | 11 (9) | 27 (10) | 7.42 | 4.72 | 2.91 |
| C(27) | 66 (6) | 63 (5) | 56 (6) | 37 (10) | -5 (10) | -26 (9) | 6.74 | 4.79 | 3.05 |
| C(28) | 71 (6) | 48 (4) | 53 (6) | 13 (9) | -1 (10) | -16 (9) | 5.82 | 4.46 | 3.32 |
| C(29) | 43 (4) | 63 (5) | 52 (6) | -10 (9) | 6 (8) | -4 (7) | 5.19 | 3.96 | 3.31 |
| O(30) | 51 (3) | 47 (3) | 45 (3) | -4 (5) | -12 (5) | 15 (5) | 4.69 | 3.48 | 3.15 |
| C(31) | 85 (7) | 37 (4) | 80 (8) | -22 (9) | 5 (12) | 11 (9) | 6.81 | 6.51 | 2.67 |
| O(32) | 53 (3) | 52 (3) | 56 (4) | 17 (6) | 12 (6) | 3 (5) | 5.17 | 4.03 | 3.50 |
| O(33) | 58 (3) | 47 (3) | 47 (3) | 12 (5) | -25 (6) | 1 (5) | 5.25 | 3.88 | 2.83 |
| C(34) | 44 (4) | 57 (5) | 60 (6) | -7 (9) | -11 (8) | 2 (7) | 5.05 | 4.37 | 3.35 |
| O(35) | 56 (3) | 50 (3) | 40 (3) | -12 (5) | 26 (5) | -5 (5) | 5.11 | 3.86 | 2.51 |

Discussion of the structure

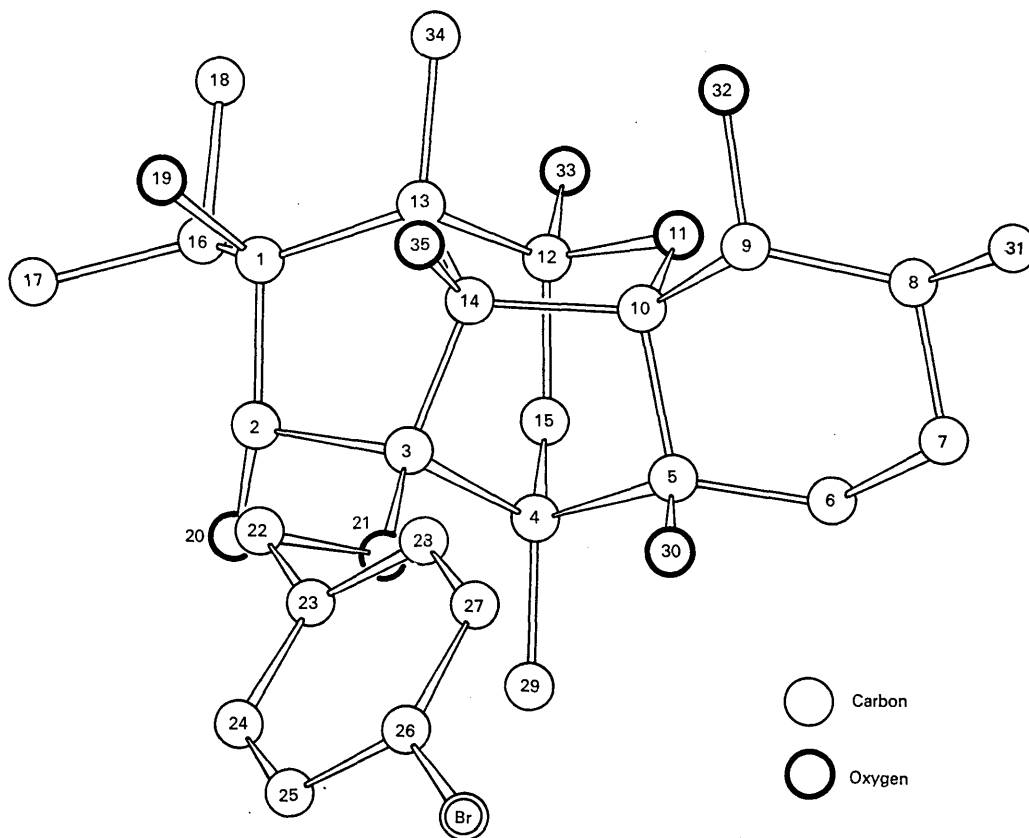
A perspective view of the molecule showing the atom numbering is given in Fig. 2.

The bond lengths are given in Table 6, and the angles are listed in Table 7. The estimated standard deviations of the bonds and angles, calculated by the expressions of Ahmed & Cruickshank (1953) and from *International Tables for X-ray Crystallography* (1959) respectively, based on the e.s.d.'s of the coordinates in Table 1, were considered to be too low because of the use of block-diagonal rather than full-matrix refinement, and also because thermal corrections were not applied and the unobserved reflexions were omitted. The e.s.d. values were therefore multiplied by two giving a range of 0.02 to 0.03 Å for the bonds and 1.2 to 2.2° for the angles. The highest e.s.d.'s were obtained for the bonds of the isopropyl group and of the benzene ring. On calculating the e.s.d. of a C(*sp*³)-C(*sp*³) bond from the scatter from the mean, a value of 0.024 Å was obtained. The observed C-C bond length variations are reasonable and their mean value (1.541 Å) agrees well with 1.537 Å given by Sutton (1965). The C(22)-C(23) bond of 1.48 Å is also in agreement with the accepted value of 1.501 Å for the C(*sp*³)-C(*sp*²) bond (Lide, 1962).

Table 6. Bond lengths

| | | | |
|-------------|--------|-------------|--------|
| C(1)-C(2) | 1.58 Å | C(16)-C(17) | 1.52 Å |
| C(1)-C(13) | 1.56 | C(16)-C(18) | 1.51 |
| C(1)-C(16) | 1.56 | C(22)-C(23) | 1.48 |
| C(2)-C(3) | 1.55 | C(23)-C(24) | 1.37 |
| C(3)-C(4) | 1.54 | C(23)-C(28) | 1.36 |
| C(3)-C(14) | 1.53 | C(24)-C(25) | 1.40 |
| C(4)-C(5) | 1.56 | C(25)-C(26) | 1.37 |
| C(4)-C(15) | 1.57 | C(26)-C(27) | 1.36 |
| C(4)-C(29) | 1.51 | C(27)-C(28) | 1.36 |
| C(5)-C(6) | 1.49 | O(11)-C(10) | 1.44 |
| C(5)-C(10) | 1.54 | O(11)-C(12) | 1.46 |
| C(6)-C(7) | 1.54 | O(19)-C(1) | 1.43 |
| C(7)-C(8) | 1.52 | O(20)-C(2) | 1.44 |
| C(8)-C(9) | 1.55 | O(21)-C(3) | 1.43 |
| C(8)-C(31) | 1.54 | O(21)-C(22) | 1.43 |
| C(9)-C(10) | 1.51 | O(30)-C(5) | 1.43 |
| C(10)-C(14) | 1.58 | O(32)-C(9) | 1.44 |
| C(12)-C(13) | 1.56 | O(33)-C(12) | 1.36 |
| C(12)-C(15) | 1.53 | O(35)-C(14) | 1.42 |
| C(13)-C(14) | 1.56 | Br — C(26) | 1.91 |
| C(13)-C(34) | 1.54 | | |

The average values for the benzene ring bonds and angles are 1.371 Å and 120.0° respectively and they compare well with those obtained by Abrahamsson & Nilsson (1966) for *p*-bromobenzyl norbormide (1.383 Å and 120.2°).

Fig. 2. Perspective view of the molecule of ryanodol *p*-bromobenzyl ether.

The O—C bonds vary from 1.42 to 1.46 Å with the exception of the O(33)—C(12) bond of 1.36 Å. We are grateful to the referee for pointing out that this bond is not significantly shorter than the equatorial carbon-oxygen bonds observed in the anomeric position in pyranose sugars (Berman, Chu & Jeffrey, 1967).

The C—H bonds of the methyl groups range from 1.05 to 1.11 Å and the average of 30 angles of these groups is 109.5°. They ranged from 104.3 to 113.9°. The O—H bond lengths are listed in Table 8.

All the six-membered rings of the skeleton are in the chair form and the five-membered rings are in the envelope form.

The mean plane of the benzene ring, referred to axes $X'=ax$, $Y'=by$ and $Z'=cz$ is $0.5841X' - 0.2186Y' - 0.7817Z' + 0.8995 = 0$. It was calculated using the procedure described by Blow (1960). The maximum distance of 0.011 Å from that plane was obtained for the C(28) atom; however, χ^2 is 2.7 and $P=40\%$ ($\nu=3$) and therefore the deviations of the six atoms of

Table 7. Bond angles

| | | | |
|------------------|--------|-------------------|--------|
| C(2)—C(1)—C(13) | 103.8° | C(9)—C(10)—O(11) | 108.0° |
| C(2)—C(1)—C(16) | 113.2 | C(9)—C(10)—C(14) | 121.0 |
| C(2)—C(1)—O(19) | 105.7 | O(11)—C(10)—C(14) | 103.6 |
| C(13)—C(1)—C(16) | 119.8 | C(10)—O(11)—C(12) | 105.6 |
| C(13)—C(1)—O(19) | 105.6 | O(11)—C(12)—C(13) | 100.3 |
| C(16)—C(1)—O(19) | 107.7 | O(11)—C(12)—C(15) | 106.1 |
| C(1)—C(2)—C(3) | 106.2 | O(11)—C(12)—O(33) | 108.6 |
| C(1)—C(2)—O(20) | 115.0 | C(13)—C(12)—C(15) | 113.7 |
| C(3)—C(2)—O(20) | 110.6 | C(13)—C(12)—O(33) | 112.7 |
| C(2)—C(3)—C(4) | 115.8 | C(15)—C(12)—O(33) | 114.2 |
| C(2)—C(3)—C(14) | 107.2 | C(1)—C(13)—C(12) | 115.3 |
| C(2)—C(3)—O(21) | 110.5 | C(1)—C(13)—C(14) | 106.3 |
| C(4)—C(3)—C(14) | 101.3 | C(1)—C(13)—C(34) | 114.4 |
| C(4)—C(3)—O(21) | 105.7 | C(12)—C(13)—C(14) | 96.8 |
| C(14)—C(3)—O(21) | 116.4 | C(12)—C(13)—C(34) | 108.6 |
| C(3)—C(4)—C(5) | 102.0 | C(14)—C(13)—C(34) | 114.3 |
| C(3)—C(4)—C(15) | 106.3 | C(3)—C(14)—C(10) | 107.3 |
| C(3)—C(4)—C(29) | 118.3 | C(3)—C(14)—C(13) | 101.6 |
| C(5)—C(4)—C(15) | 106.7 | C(3)—C(14)—O(35) | 113.1 |
| C(5)—C(4)—C(29) | 111.7 | C(10)—C(14)—C(13) | 104.9 |
| C(15)—C(4)—C(29) | 111.0 | C(10)—C(14)—O(35) | 113.5 |
| C(4)—C(5)—C(6) | 121.2 | C(13)—C(14)—O(35) | 115.4 |
| C(4)—C(5)—C(10) | 100.4 | C(4)—C(15)—C(12) | 118.1 |
| C(4)—C(5)—O(30) | 110.2 | C(1)—C(16)—C(17) | 120.9 |
| C(6)—C(5)—C(10) | 110.4 | C(1)—C(16)—C(18) | 118.3 |
| C(6)—C(5)—O(30) | 105.3 | C(17)—C(16)—C(18) | 118.8 |
| C(10)—C(5)—O(30) | 109.1 | C(3)—O(21)—C(22) | 119.7 |
| C(5)—C(6)—C(7) | 108.4 | O(21)—C(22)—C(23) | 121.5 |
| C(6)—C(7)—C(8) | 115.4 | C(22)—C(23)—C(24) | 118.8 |
| C(7)—C(8)—C(9) | 114.6 | C(22)—C(23)—C(28) | 122.5 |
| C(7)—C(8)—C(31) | 109.9 | C(24)—C(23)—C(28) | 111.5 |
| C(9)—C(8)—C(31) | 109.4 | C(23)—C(24)—C(25) | 110.4 |
| C(8)—C(9)—C(10) | 110.0 | C(24)—C(25)—C(26) | 113.1 |
| C(8)—C(9)—O(32) | 110.5 | Br—C(26)—C(25) | 108.8 |
| C(10)—C(9)—O(32) | 110.3 | Br—C(26)—C(27) | 121.7 |
| C(5)—C(10)—C(9) | 111.9 | C(25)—C(26)—C(27) | 107.6 |
| C(5)—C(10)—O(11) | 107.7 | C(26)—C(27)—C(28) | 121.0 |
| C(5)—C(10)—C(14) | 103.8 | C(23)—C(28)—C(27) | 121.0 |

Table 8. Examination of the shortest O...O contacts

| Intramolecular | \angle C—O—H | O—H | H...O | O...O | \angle O—H...O |
|----------------------|----------------|--------|--------|--------|------------------|
| O(19)—H(191)...O(35) | 123.4° | 1.01 Å | 2.52 Å | 3.11 Å | 116.5° |
| O(20)—H(201)...O(19) | 77.8 | 1.11 | 2.96 | 3.56 | 114.5 |
| O(33)—H(331)...O(11) | 109.1 | 1.00 | 2.41 | 2.29* | 71.3 |
| O(35)—H(351)...O(32) | 107.1 | 0.97 | 2.64 | 3.06 | 106.9 |
| Intermolecular | | | | | |
| O(20)—H(201)...O(30) | 77.8 | 1.11 | 2.68 | 2.78 | 83.0 |
| O(30)—H(301)...O(20) | 117.9 | 0.99 | 2.54 | 2.78 | 93.5 |
| O(32)—H(321)...O(35) | 115.3 | 0.99 | 2.58 | 3.45 | 145.6 |
| O(33)—H(331)...O(32) | 109.1 | 1.00 | 1.74 | 2.69 | 159.0 |
| O(35)—H(351)...O(11) | 107.1 | 0.97 | 2.47 | 3.07 | 119.9 |

* The unusually short O(33)...O(11) distance is due to the fact that the epoxy oxygen atom O(11) is separated only by one carbon atom from O(33) of the hydroxyl group.

the benzene ring from the plane are not significant. The C(22) and Br atoms were found to be situated on the opposite sides of the plane, at distances of 0.03 and 0.04 Å respectively and the C(22)–C(23) and C(26)–Br bonds make angles of only 1.2° with the benzene ring.

The molecules related by the twofold screw axes parallel to the *c* axis form hydrogen-bonded spirals. The packing of the molecules along the *c* axis can be seen in Fig. 3, and Fig. 4 shows three molecules of the spiral. The hydrogen bonding joining these molecules is between the atoms O(33) and O(32), which are 2.69 Å apart. The H(331)···O(32) distance was found to be 1.74 Å and \angle O(33)–H(331)···O(32) is 159.0°. This is the only hydrogen bond in the structure. All other hydrogen atoms of OH groups were found to make at least one intra- or intermolecular H···O contact,

which does not deviate significantly from the sum of the van der Waals radii. The hydrogen atoms are directed towards the neighbouring oxygen atoms optimizing the H···O interactions. The shortest O···O distances are examined in Table 8.

The O···C intermolecular approaches are normal, the closest occurring between the molecules of the hydrogen-bonded spiral, and they are: O(33)···C(9), 3.26 Å and O(19)···C(31), 3.30 Å. The smallest C···C distance is 3.56 Å.

The closest contacts of the bromine atom are to the atoms O(33) and C(22). The Br···O(33) distance is 3.69 Å with H(331)···Br of 3.47 Å and Br···C(22) is 3.66 Å, the atoms H(221) and H(222) being situated at 3.10 and 3.30 Å from the bromine atom. All other bromine to carbon distances were found to be above 3.8 Å.

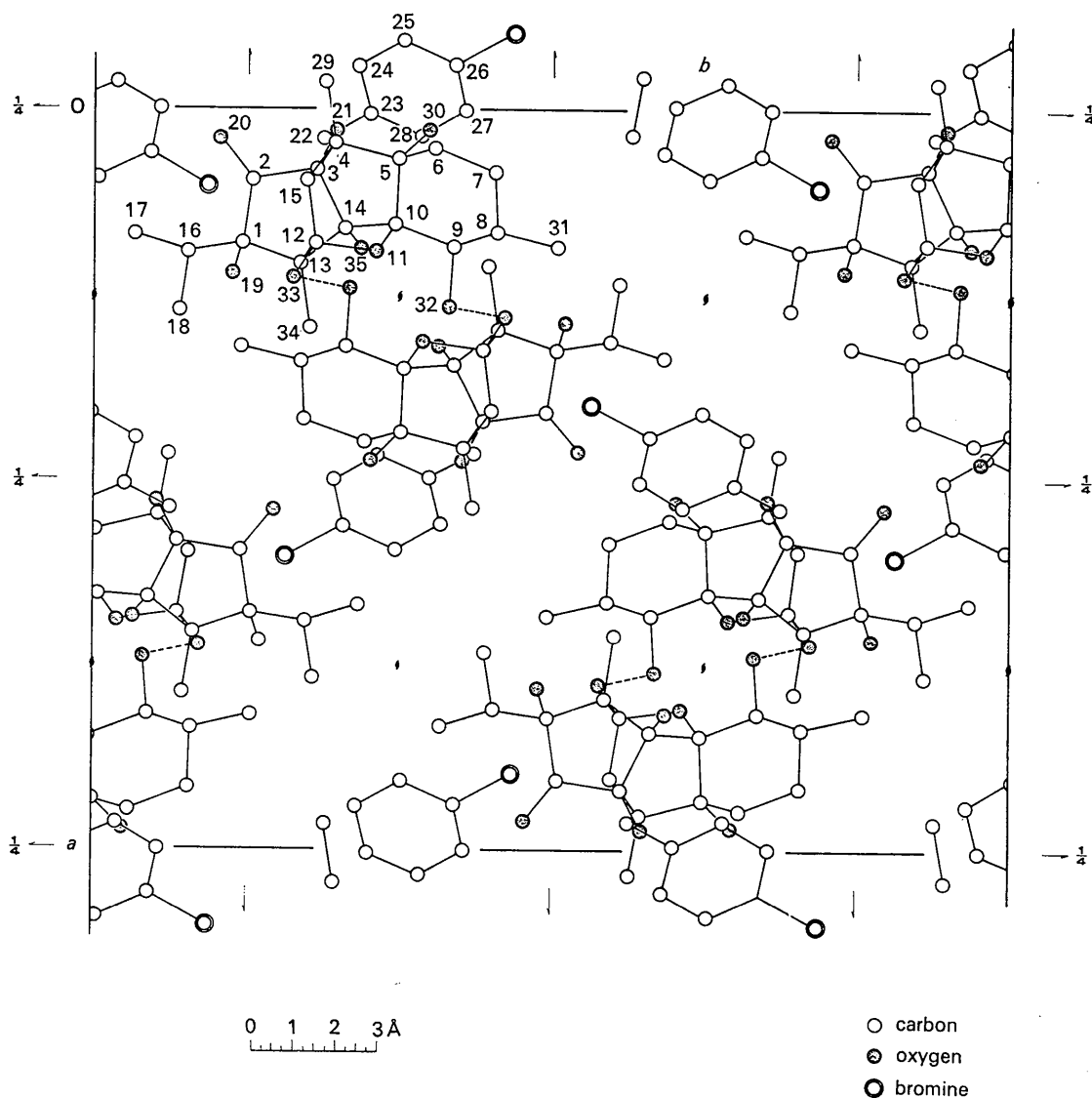


Fig. 3. Packing of the molecules in the (001) projection.

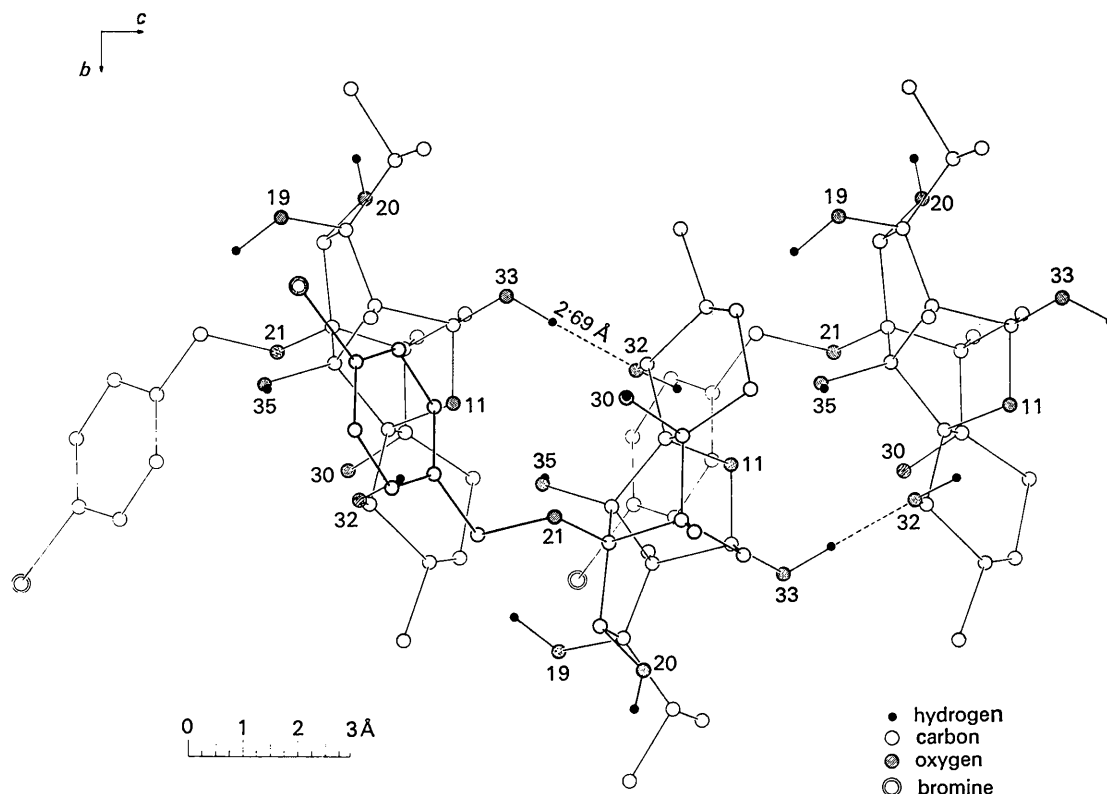


Fig. 4. Projection of three molecules along the a axis showing the hydrogen bonds.

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